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## Remember!

- REMEMBER...The Bronsted-Lowry Theory States
  - Bases accept  $H^+$  and acids donate  $H^+$
  - A conjugate base is what remains after the acid gives up  $H^+$ .
  - A conjugate acid is what is formed when a base accepts a proton.

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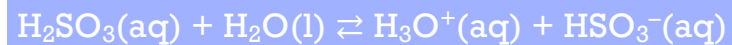
1. Which compound will donate a proton? \_\_\_\_\_
2. What does the compound look like after it donates a proton? \_\_\_\_\_
3. Which compound will accept a proton? \_\_\_\_\_
4. What does the compound look like after it accepts a proton? \_\_\_\_\_
5. Identify the acid, base, conjugate acid and conjugate base below.

Acid	Conjugate Base	Base	Conjugate Acid

4



5



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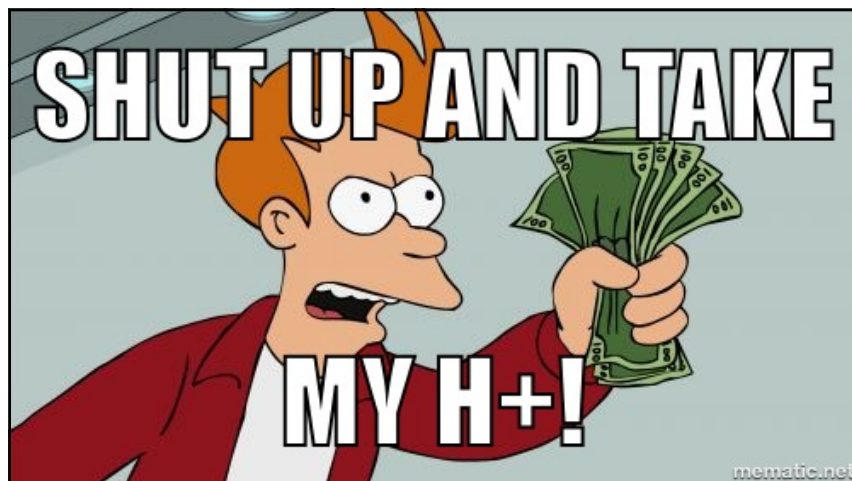
13



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14



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