

Name: KEY

Off. Class: \_\_\_\_\_

WORK @ END

Per: \_\_\_\_\_

Date: \_\_\_\_\_

Teacher: \_\_\_\_\_

### Unit 8 Thermochemistry Practice Test

Chemistry

1. What is the total number of kilojoules required to boil 100 grams of water at 100°C and 1 atmosphere?

- A) 33.4 kJ
- C) 22.6 kJ

B) 334 kJ  $q = m \cdot H_v$   
 D) 226 kJ  $q = 100 \text{ g} \cdot 2260 \text{ J/g}$

2. What occurs when a substance melts?

- ~~A) It changes from liquid to solid, and heat is absorbed.~~
- ~~B) It changes from liquid to solid, and heat is released.~~
- C) It changes from solid to liquid, and heat is released.

D) It changes from solid to liquid, and heat is absorbed.

heat added

3. What is the minimum number of kilojoules needed to change 40.0 grams of water at 100°C to steam at the same temperature and pressure?

- A) 90.4 kJ
- C) 1,810 kJ

B) 2.26 kJ  $q = 40 \text{ g} \cdot 2260 \text{ J/g}$   
 D) .400 kJ  $= 90400 \text{ J}$

$kJ = 90.4$

4. Which statement describes the transfer of heat energy that occurs when an ice cube is added to an insulated container with 100 milliliters of water at 25°C?

- ~~A) Both the ice cube and the water lose heat energy.~~
- ~~B) Both the ice cube and the water gain heat energy.~~
- C) The ice cube gains heat energy and the water loses heat energy.
- D) The ice cube loses heat energy and the water gains heat energy.

heat moves warm to cold - water → ice

5. The heat of vaporization of a liquid is 1,340 Joules per gram. What is the minimum number of Joules needed to change 40.0 grams of the liquid to vapor at the boiling point?

- A) 33.5
- C) 53,600

B) 1,340  $m = 40.0$   
 D) 3,280  $q = ?$   
 $q = (1340)(40)$

6. When 420 Joules of heat energy is added to 10. grams of water at 20.°C, the final temperature of the water will be Find  $\Delta T$  first - then  $\Delta T = T_f - T_i$

- A) 40.°C
- C) 10.°C

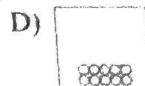
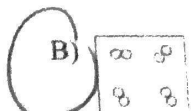
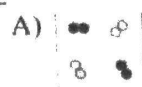
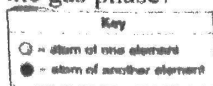
B) 30.°C  
D) 100.°C

7. Which change of phase is exothermic?

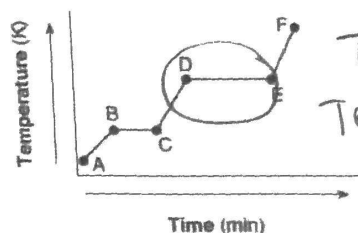
- A)  $\text{H}_2\text{O}(s) \rightarrow \text{H}_2\text{O}(g)$
- B)  $\text{CO}_2(s) \rightarrow \text{CO}_2(l)$
- C)  $\text{NH}_3(l) \rightarrow \text{NH}_3(g)$
- D)  $\text{H}_2\text{S}(g) \rightarrow \text{H}_2\text{S}(l)$

heat added endo  
heat released exo

8. Which particle diagram represents one substance in the gas phase?



9. The heating curve below represents a sample of a substance starting as a solid below its melting point and being heated over a period of time.

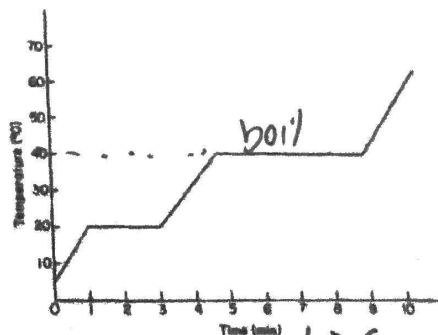


Temp = KE  
Temp const. KE = const.  
DEA

Which statement describes the energy of the particles in this sample during interval DE?

- A) Both potential energy and average kinetic energy increase.
- B) Both potential energy and average kinetic energy decrease.
- C) Potential energy remains the same and average kinetic energy increases.
- D) Potential energy increases and average kinetic energy remains the same.

10. The graph below represents changes of state for an unknown substance.



What is the boiling temperature of the substance?

- A) 0°C
- B) 40°C
- C) 20°C
- D) 70°C

WORK ON BACK

11. Which statement explains why H<sub>2</sub>O has a higher boiling point than N<sub>2</sub>?  $\uparrow MF \uparrow BP$
- A) H<sub>2</sub>O has greater molar mass than N<sub>2</sub>.  
 B) H<sub>2</sub>O has less molar mass than N<sub>2</sub>.  
 C) H<sub>2</sub>O has weaker intermolecular forces than N<sub>2</sub>.  
 D) H<sub>2</sub>O has stronger intermolecular forces than N<sub>2</sub>

12. What is the total number of kilojoules of heat energy absorbed when the temperature of 200 grams of water is raised from 10°C to 40°C?

A) 0.126 kJ  
 B) 0.840 kJ  
 C) 33.6 kJ  
 D) 25.2 kJ

*q = mcΔT*

13. Which equation represents sublimation?  $S \rightarrow G$  directly
- A) I<sub>2</sub>(s) → I<sub>2</sub>(g)  
 B) I<sub>2</sub>(s) → I<sub>2</sub>(l)  
 C) I<sub>2</sub>(l) → I<sub>2</sub>(s)  
 D) I<sub>2</sub>(l) → I<sub>2</sub>(g)

14. A sample of water is heated from 10.0°C to 15.0°C by the addition of 126 Joules of heat. What is the mass of the water?

A) 5.00 g  
 B) 150.0 g  
 C) 6.00 g  
 D) 30.0 g

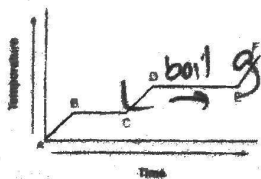
*q = mcΔT*

15. The temperature of 50.0 grams of water was raised to 50.0°C by the addition of 4200 Joules of heat energy. What was the initial temperature of the water?

A) 10.0°C  
 B) 20.0°C  
 C) 30.0°C  
 D) 60.0°C

*q = mcΔT*

16. The graph below represents the uniform heating of a substance from the solid to the gas phase.



Which line segment of the graph represents boiling?

A) AB B) BC C) CD D) DE

17. How many Joules of heat energy are absorbed in raising the temperature of 10. grams of water from 5.0°C to 20.°C?

A)  $6.3 \times 10^2$   
 B)  $2.1 \times 10^2$   
 C)  $8.4 \times 10^2$   
 D)  $1.1 \times 10^3$

*q = mcΔT*

18. A student observing the behavior of paradichlorobenzene first heats 10 grams of the substance in a hot water bath until it is completely liquefied. The following data are recorded as paradichlorobenzene cools.

DATA TABLE

Time (minutes)	Temperature (°C)
0	65
1	58
2	52
3	53
4	53
5	53
6	53
7	53
8	51
9	47
10	42

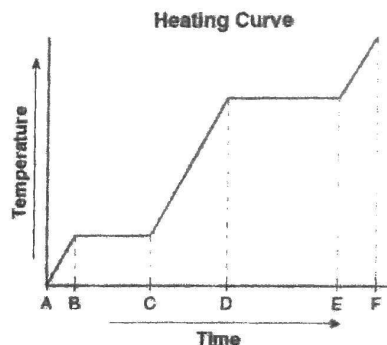
Temp is constant when freezing (phase change)

53 is the only constant temp

What is the freezing point of paradichlorobenzene?

A) 42°C B) 53°C C) 58°C D) 65°C

19. Given the diagram representing a heating curve for a substance:



During which time interval is the average kinetic energy of the particles of the substance constant while the potential energy of the particles increases?

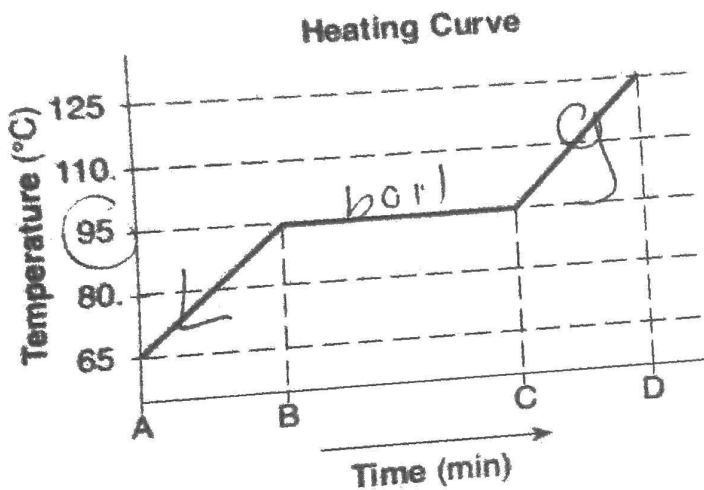
A) CD B) BC C) DF D) AC

↳ flat section only

work on back

Base your answers to questions 20 and 21 on the information below and on your knowledge of chemistry.

A sample of a substance is a liquid at 65°C. The sample is heated uniformly to 125°C. The heating curve for the sample at standard pressure is shown below.

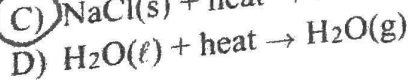
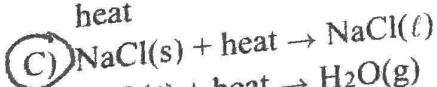
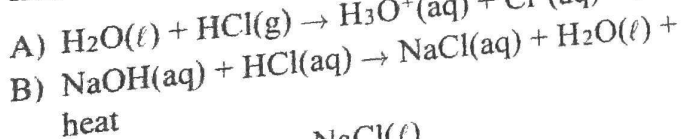
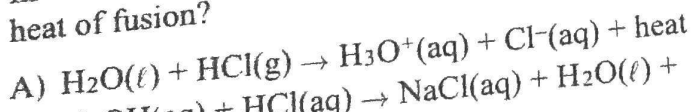


20. Determine the boiling point of the sample at standard pressure.

21. State what happens to the potential energy of the particles of the sample during time interval BC.

*KE = constant PE increases*

22. In which equation does the term "heat" represent heat of fusion?



*- used when something is going from solid or liq → solid*

25. Base your answer to the following question on the information below.

In a laboratory, a student makes a solution by completely dissolving 80.0 grams of  $\text{KNO}_3(\text{s})$  in 100.0 grams of hot water. The resulting solution has a temperature of 60.°C. The room temperature in the laboratory is 22°C.

Describe the direction of heat flow between the solution made by the student and the air in the laboratory.

*heat flow from solution to air*

*(heat moves from warm to cool)*



# 6, 12, 14, 15, 17

*[Handwritten scribble]*

#6  $q = mc\Delta T$   $\Delta T = T_f - T_i$  (Find  $\Delta T$  first then  $T_f$ )

$q = 420 \text{ J}$

$m = 10 \text{ g}$

$c = 4.18 \text{ J/gC}$  (value on RT)

$\Delta T =$

$T_f =$

$T_i = 20^\circ\text{C}$

$420 \text{ J} = (10 \text{ g}) (4.18 \text{ J/gC}) (\Delta T)$

$\frac{420 \text{ J}}{41.8} = \frac{41.8 (\Delta T)}{41.8}$

$10 = \Delta T$

$\Delta T = T_f - T_i$

$10 = T_f - 20$   
 $+20$

$T_f = 30^\circ\text{C}$

12.  $q = mc\Delta T$

$q = ?$

$m = 200 \text{ g}$

$c = 4.18 \text{ J/gC}$  (RT)

$\Delta T = 30^\circ\text{C}$

$T_f = 40^\circ\text{C}$

$T_i = 10^\circ\text{C}$

$q = (200 \text{ g}) (4.18 \text{ J/gC}) (30^\circ\text{C})$

$q = 25080 \text{ J}$

$q = 25.08 \text{ kJ}$

25.2 choice B  
is essentially the  
same thing  
Make sure to convert to kJ

14.  $q = mc\Delta T$

$q = 126 \text{ J}$

$m = ?$

$c = 4.18 \text{ J/gC}$

$\Delta T = 5^\circ\text{C}$

$T_f = 15^\circ\text{C}$

$T_i = 10^\circ\text{C}$

$126 = m (4.18) (5)$

$\frac{126}{20.9} = \frac{m (20.9)}{20.9}$

$m = 6.02 \text{ g}$  choice C

#15.  $q = mc\Delta T$

$q = 4200 \text{ J}$

$m = 50.0 \text{ g}$

$c = 4.18 \text{ J/g}\cdot\text{C}$

$\Delta T =$

$T_f = 50.0^\circ\text{C}$

$T_i =$

$$4200 = (50)(4.18)\Delta T$$

$$\frac{4200}{209} = \frac{209}{209}\Delta T$$

$\Delta T = 20.10$

$\Delta T = T_f - T_i$

$$20 = 50 - T_i$$

$$\begin{array}{r} +T_i \\ 20 + T_i = 50 \\ -20 \quad -20 \end{array}$$

$T_i = 30^\circ\text{C}$

choice C

#17.  $q = mc\Delta T$

$q = ?$

$m = 10 \text{ g}$

$c = 4.18 \text{ J/g}\cdot\text{C}$

$\Delta T = 15.0^\circ\text{C}$

$T_f = 20^\circ\text{C}$

$T_i = 5^\circ\text{C}$

$$q = (10)(4.18)(15)$$

$q = 627 \text{ J}$

$= 6.27 \times 10^2 \text{ J}$

choice A

$6.30 \times 10^2 = 630 \text{ J}$