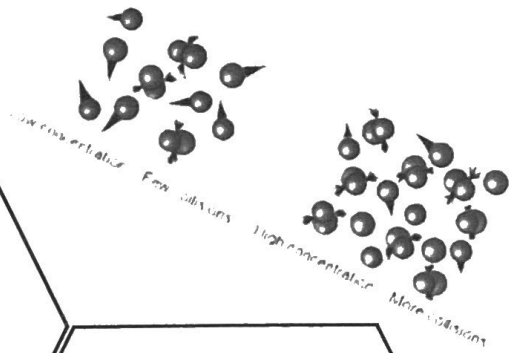
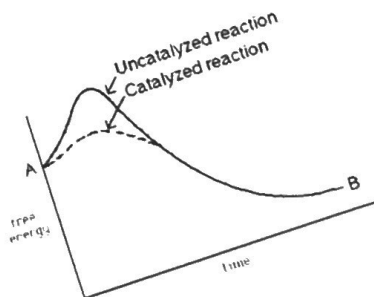


Factors Affecting the Rates of Reaction



Type of Reactant

- IONIC substances react faster. They easily break into ions when you dissolve them

- COVALENT substances react slower - Requires more energy/time to break bonds

Catalyst

- increases the rate of reaction by lowering the activation energy

- Is NOT consumed in the reaction

Concentration

- increase concentration increase the reaction rate (speed)

- More particles increase chance of effective collisions

Background Info:

- There are SIX Factors that affect the rate of reaction by changing the number of effective collisions that take place between particles
- The more effective collisions, the faster the reaction

Surface Area

- increase surface area increase the reaction rate (speed)

- Due to more exposed particles that can react (more effective collisions)

Temperature

- increase temperature increase the reaction rate (speed)

- Increases the number effective collisions

- Reactants have more energy when colliding

Pressure

- increase pressure increase the reaction rate (speed)

affect gases only

- Due to an increase in concentration

