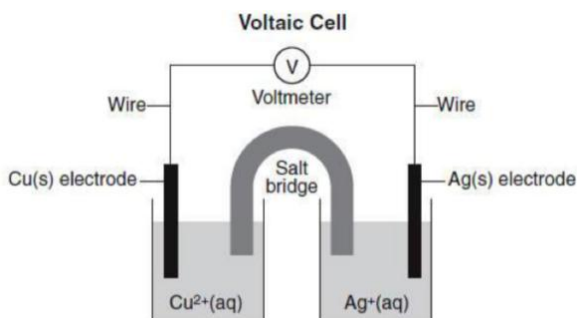


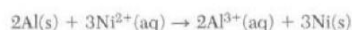
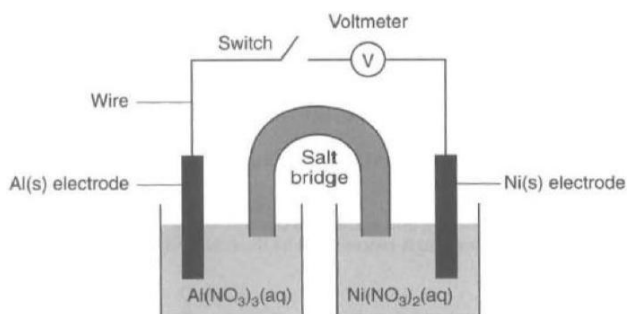
Name: \_\_\_\_\_ Official Class: \_\_\_\_\_ Date: \_\_\_\_\_  
 Teacher: \_\_\_\_\_ Period: \_\_\_\_\_ Class: \_\_\_\_\_

## More Voltaic Cell Practice

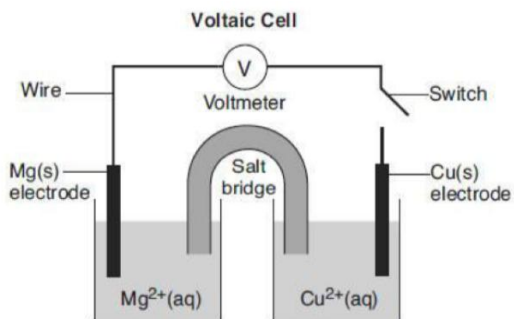
**Directions:** Answer the following questions based on the diagrams and your knowledge of chemistry.



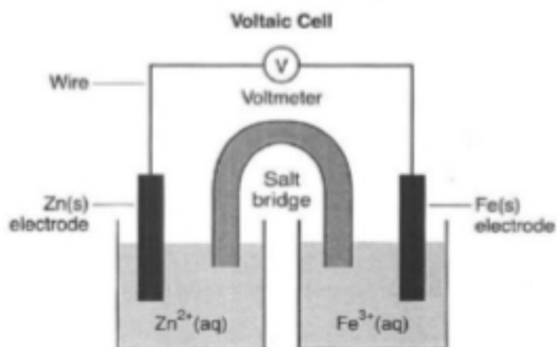
- Anode: \_\_\_\_\_
- Cathode: \_\_\_\_\_
- Direction of e<sup>-</sup>: \_\_\_\_\_
- \_\_\_\_\_ increases in mass
- \_\_\_\_\_ decreases in mass
- Oxidation half reaction: \_\_\_\_\_
- Reduction half reaction: \_\_\_\_\_



- Anode: \_\_\_\_\_
- Cathode: \_\_\_\_\_
- Direction of e<sup>-</sup>: \_\_\_\_\_
- \_\_\_\_\_ increases in mass
- \_\_\_\_\_ decreases in mass
- Oxidation half reaction: \_\_\_\_\_
- Reduction half reaction: \_\_\_\_\_



- Anode: \_\_\_\_\_
- Cathode: \_\_\_\_\_
- Direction of e<sup>-</sup>: \_\_\_\_\_
- \_\_\_\_\_ increases in mass
- \_\_\_\_\_ decreases in mass
- Oxidation half reaction: \_\_\_\_\_
- Reduction half reaction: \_\_\_\_\_



- Anode: \_\_\_\_\_
- Cathode: \_\_\_\_\_
- Direction of e<sup>-</sup>: \_\_\_\_\_
- \_\_\_\_\_ increases in mass
- \_\_\_\_\_ decreases in mass
- Oxidation half reaction: \_\_\_\_\_
- Reduction half reaction: \_\_\_\_\_

- In terms of atoms and ions, why is the mass of the anode decreasing?
- In terms of atoms and ions, why is the mass of the cathode increasing?