

1. by definition mixtures have varying compositions think fruitsalad you can have the same thing w/ different amounts of each part.

2. salt is dissolved & sand is insoluble when you put through a filter, sand (the insoluble thing) stays behind and the salt (soluble) passes through. evaporating the water leaves the salt behind.

3. the more solute added ↑ BP ↓ FP

4. (aq) = something in water solvent = almost always water (does the dissolving) solute = gets dissolved

6. ↑ temp ↑ solid solubility look @ Table G pressure only affects gases

7. HCl is a gas and ↓ solubility ↑ temp TABLE G

8. $KCl \rightarrow K^+ + Cl^-$
 $K_2SO_4 \rightarrow K^+ + K^+ + SO_4^{2-}$
 $K_3PO_4 \rightarrow K^+ + K^+ + K^+ + PO_4^{3-}$
 $KNO_3 \rightarrow K^+ + NO_3^-$

- Distillation of crude oil from various parts of the world yields different percentages of hydrocarbons. Which statement explains these different percentages?
 A) Hydrocarbons are organic compounds
 B) Each component in a mixture has a different solubility in water.
 C) The carbons in hydrocarbons may be bonded in chains or rings.
 (D) **The proportions of components in a mixture can vary.**
- A mixture consists of sand and an aqueous salt solution. Which procedure can be used to separate the sand, salt, and water from each other?
 A) Evaporate the water, then filter out the sand.
 B) Evaporate the water, then filter out the salt.
 (C) **Filter out the sand, then evaporate the water.**
 D) Filter out the salt, then evaporate the water.
- Compared to a 2.0 M aqueous solution of NaCl at 1 atmosphere, a 3.0 M aqueous solution of NaCl at 1 atmosphere has a
 A) higher boiling point and a higher freezing point
 B) lower boiling point and a lower freezing point
 (C) **higher boiling point and a lower freezing point**
 D) lower boiling point and a higher freezing point
- In an aqueous solution of potassium chloride, the solute is
 A) K (B) **KCl** C) Cl D) H₂O
- When KCl(s) is dissolved in water, the resulting solution is classified as a
 A) heterogeneous mixture
 B) homogeneous compound
 C) heterogeneous compound
 (D) **homogeneous mixture**
- The solubility of KClO₃(s) in water increases as the
 (A) **temperature of the solution increases**
 B) pressure on the solution increases
 C) pressure on the solution decreases
 D) temperature of the solution decreases
- At standard pressure, which substance becomes less soluble in water as temperature increases from 10.°C to 80.°C?
 A) NH₄Cl (B) **HCl** C) NaCl D) KCl
- Which solution has the highest boiling point at standard pressure?
 A) 0.10 M KCl(aq) 2 pieces B) 0.10 M K₂SO₄(aq) 3 pieces
 (C) **0.10 M K₃PO₄(aq)** 4 D) 0.10 M KNO₃(aq) 2 pieces
- A bottle of rubbing alcohol contains both 2-propanol and water. These liquids can be separated by the process of distillation because the 2-propanol and water
 A) have combined chemically and retain their different boiling points
 B) have combined physically and have the same boiling point
 (C) **have combined physically and retain their different boiling points**
 D) have combined chemically and have the same boiling point
- Differences in which property allow the separation of a sample of sand and seawater by filtration?
 (A) **particle size** B) concentration of ions
 C) volume of sample D) mass of sample
- When 5 grams of KCl are dissolved in 50. grams of water at 25°C, the resulting mixture can be described as
 (A) **homogeneous and unsaturated**
 B) heterogeneous and supersaturated
 C) heterogeneous and unsaturated
 D) homogeneous and supersaturated

on Table G the point is UNDER the curve = UNSATURATED
 Sat & UNSat = homogeneous
 Supersat = heterogeneous

- Based on Table F, which equation represents a saturated solution having the lowest concentration of Cl⁻ ions?
 A) $KCl(s) = K^+(aq) + Cl^-(aq)$
 B) $NH_4Cl(s) = NH_4^+(aq) + Cl^-(aq)$
 C) $NaCl(s) = Na^+(aq) + Cl^-(aq)$
 (D) **$AgCl(s) = Ag^+(aq) + Cl^-(aq)$** ← exception on Table F
- A 1-gram sample of a compound is added to 100 grams of H₂O(l) and the resulting mixture is then thoroughly stirred. Some of the compound is then separated from the mixture by filtration. Based on Table F, the compound could be
 A) CaCl₂ (B) **AgCl** C) NiCl₂ D) NaCl
- Paper chromatography can separate the components of a mixture of colored dyes because the components have differences in
 A) thermal conductivity B) decay mode
 C) ionization energy (D) **molecular polarity**
- According to Table G, which substance forms an unsaturated solution when 80. grams of the substance are stirred into 100. grams of H₂O at 10.°C?
 A) KNO₃ B) NH₃ C) NaCl (D) **KI**
- What is the concentration of a solution, in parts per million, if 0.02 gram of Na₃PO₄ is dissolved in 1000 grams of water?
 A) 2 ppm B) 0.2 ppm
 C) 0.02 ppm (D) **20 ppm**
- The solubility of KCl(s) in water depends on the
 A) size of the KCl sample
 (B) **temperature of the water**
 C) pressure on the solution
 D) rate of stirring
- What is the concentration of an aqueous solution that contains 1.5 moles of NaCl in 500 milliliters of this solution?
 A) 7.5 M B) 0.30 M C) 0.75 M (D) **3.0 M**
- The molarity of an aqueous solution of NaCl is defined as the
 A) grams of NaCl per liter of solution
 B) moles of NaCl per liter of water
 C) grams of NaCl per liter of water
 (D) **moles of NaCl per liter of solution**
- Which substance is most soluble in water?
 F) A) CaCO₃ B) Cu(OH)₂
 F) C) Ag₂SO₄ (D) **(NH₄)₃PO₄**

I = insoluble
 S = soluble
 Determine the Solubility of each thing - one will be different - that is your answer

16. Solute = 0.02 g
 Solution = 1000.02 g
 $\frac{0.02g}{1000.02g} \times 10^6 = 20 \text{ ppm}$
 = 19.9996 ↑

18. $M = \frac{n}{L}$
 M = ?
 n = 1.5 mol
 L = 500 mL → 0.5 L
 $M = \frac{1.5 \text{ mol}}{0.5 \text{ L}} = 3 \text{ M}$
 MUST USE LITERS
 mL → L
 more decimal 3 spots to left.

10. Sand is too large to pass through filter paper

$M = \frac{n}{L}$ straight definition formula on Table T

UNSATURATED = under the curve

DEFINITION - JUST KNOW IT

chromatography = separates by polarity

insoluble things are separated by filtration - AgCl is an exception and is INSOLUBLE

lowest concentration = insoluble

Base your answers to questions 21 and 22 on the information below and on your knowledge of chemistry.

A student prepares two 141-gram mixtures, A and B. Each mixture consists of NH_4Cl , sand, and H_2O at 15°C . Both mixtures are thoroughly stirred and allowed to stand. The mass of each component used to make the mixtures is listed in the data table below.

Mass of the Components in Each Mixture

Component	Mixture A (g)	Mixture B (g)
NH_4Cl	40.	10.
sand	1	31
H_2O	100.	100.

Soluble →
insoluble →

21. Describe *one* property of sand that would enable the student to separate the sand from the other components in mixture B.

Sand is insoluble - stays solid
- too big to fit through filter

22. Determine the temperature at which all of the NH_4Cl in mixture A dissolves to form a saturated solution.

23°C - 26°C - any # in that range is acceptable
- look @ Table G @ NH_4Cl line - @ 40g - when are you on the line?

23. Base your answer to the following question on the information below and on your knowledge of chemistry.

Some compounds of silver are listed with their chemical formulas in the table below.

Silver Compounds

Name	Chemical Formula
silver carbonate	Ag_2CO_3
silver chlorate	AgClO_3
silver chloride	AgCl
silver sulfate	Ag_2SO_4

insoluble
soluble
insoluble
insoluble

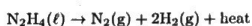
Identify the silver compound in the table that is most soluble in water.

(one thing will be soluble & the others will be insoluble) TABLE F

AgClO_3

24. Base your answer to the following question on the information below and on your knowledge of chemistry.

Hydrazine, N_2H_4 , is a compound that is very soluble in water and has a boiling point of 113°C at standard pressure. Unlike water, hydrazine is very reactive and is sometimes used as a fuel for small rockets. One hydrazine reaction producing gaseous products is represented by the balanced equation below.



Explain, in terms of molecular polarity, why N_2H_4 is very soluble in water.

Water & N_2H_4 are both polar

Remember: Like dissolves like, but that's not enough for credit you need to correctly

25. Base your answer to the following question on the information below.

A total of 1.4 moles of sodium nitrate is dissolved in enough water to make 2.0 liters of an aqueous solution. The gram formula mass of sodium nitrate is 85 grams per mole.

Determine the molarity of the solution.

mention polarity of the 2 compounds

Steps
+ find molar mass first

$$M = \frac{n}{L}$$

$$M = ?$$

$$n = 1.4 \text{ mol}$$

$$L = 2.0 \text{ L}$$

$$M = \frac{1.4 \text{ mol}}{2.0 \text{ L}} = 0.7 \text{ mol/L}$$

0.7 M