Name	:KEY	_	
Teach	er:	Per:	Date:
	Introduction	Official Class:	Chemistry
Introduction to Heating and Cooling Curves  YOYO: Draw out 8 particles in each box for the			
YOYO: Draw out 8 particles in each box for the given phase. Fill in the arrows with words in the word box below. I highly recommend using a pencil in case you make a mistake.			
== odbe you make a mistake.			
Word Box  Meting * Vaporization * Sublimation * Deposition * Freezing * Condensation * Removed * Added			
- Condensation * Removed * Added			
	Heat must be	added to go from a solid to a	
		to go from a solid to a	a gas
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	VY		V
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			47
	Solid	Liquid boiling	
1.	melting	evaporat	O Gas (
	Tricting	Vaporizi	ation 0
	,		$\neg \lor \mid \lor  \lor  \mid$
-	1/		0.0
	SEQ Freeziv		
-	ACC THE STATE OF T	08000 C Conden	sotion o
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			7,00
			1 San
	1		Z
	N N		
	Heat must be	removed to go from a gas to a	solid
3 to a solid			
Heat vs. Temperature			
Heat, heat energy, or thermal energy			
The state of the s			
The same of constant that The same of the			
two samples of matter because of their half lows in temperatures			
o Thermal energy transferred from a hotter system to a conter system			
that are in contact			
Temperature			
Adding heat will 10000000			
•	will the tem		heat
_			
A heating curve is a graph that shows the temperature of a substance plotted against the amount of operations.			
** ** 0	energy.		

Unit 8: Thermochemistry

Date: Name: Chemistry Official Class: Teacher: temperatue. Colling Condendation 1 PE/ prosecreme - particle diagram potential every Erretic every (sound & higher) melting 7 PE heat removed. heat added Grizzant 988 SCALINA OXTON Heating Curve of Water time JY L VKE Cooling Curve 5 0 -PE 0 TIPE vaporization, sna heating curv = TKE & constant PE Rigul dyas constant KE & TIPE condens artis 0000 0 0 **Unit 8: Thermochemistry** 0